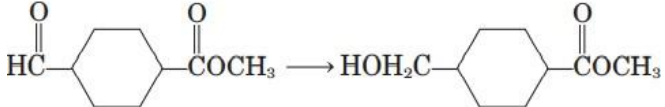
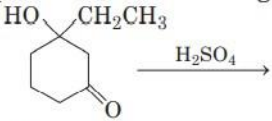
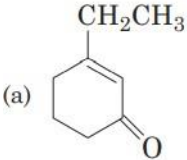
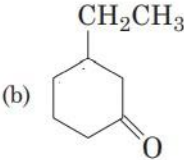
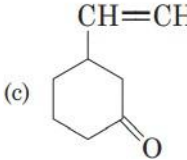
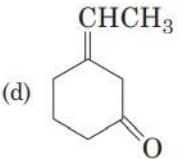
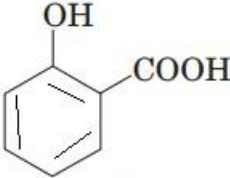
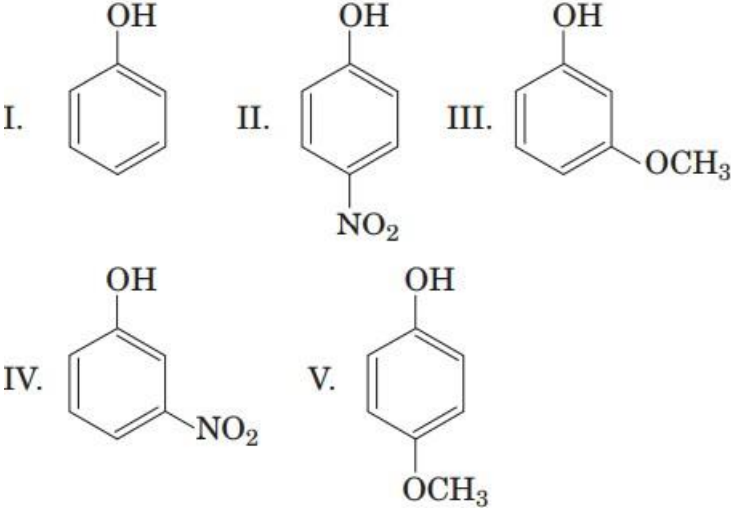




**MULTIPLE CHOICE QUESTIONS (1 MARK EACH)**

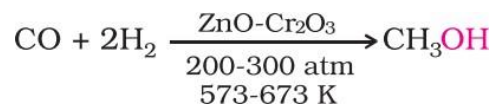
1	<p><i>Iso-propyl alcohol is</i> <span style="float: right;">(NCERT)</span></p> <p>(a) <math>\text{CH}_2-\underset{\text{CH}_3}{\text{CH}}-\text{CH}_2-\text{OH}</math> (b) <math>\begin{array}{l} \text{CH}_3 \\ \diagdown \\ \text{H}_3\text{C}-\text{C}-\text{OH} \\ \diagup \\ \text{H}_3\text{C} \end{array}</math></p> <p>(c) <math>\text{CH}_3-\underset{\text{CH}_3}{\text{CH}}-\text{OH}</math> (d) <math>\text{CH}_3-\underset{\text{CH}_2-\text{CH}_3}{\text{CH}}-\text{OH}</math></p>
2	<p>The major product of the following reaction is</p> <p><math>\text{CH}_3\text{CH}=\text{CHCO}_2\text{CH}_3 \xrightarrow{\text{LiAlH}_4} \quad (2019)</math></p> <p>(a) <math>\text{CH}_3\text{CH}=\text{CHCH}_2\text{OH}</math>          (b) <math>\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}</math>          (c) <math>\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{CH}_3</math>          (d) <math>\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}</math></p>
3	<p>The reduction,</p> <p></p> <p>can be achieved by using</p> <p>(a) <math>\text{NaBH}_4</math> (b) <math>\text{LiAlH}_4</math>          (c) <math>\text{CuO} \cdot \text{CuCN}_2\text{O}_4</math> (d) None of these</p>

4	<p style="text-align: center;"> <math display="block">\begin{array}{c} \text{OH} \\   \\ \text{CH}_3\text{CH}_2-\text{C}-\text{CH}_3 \\   \\ \text{Ph} \end{array}</math>           cannot be prepared by           <span style="float: right;">(2019)</span> </p> <p>(a) <math>\text{CH}_3\text{CH}_2\text{COCH}_3 + \text{PhMgX}</math>            (b) <math>\text{PhCOCH}_3 + \text{CH}_3\text{CH}_2\text{MgX}</math>            (c) <math>\text{PhCOCH}_2\text{CH}_3 + \text{CH}_3\text{MgX}</math>            (d) <math>\text{HCHO} + \text{PhCH}(\text{CH}_3)\text{CH}_2\text{MgX}</math></p>
5	<p>In the following reaction, identify X.            Methyl magnesium bromide + X <math>\longrightarrow</math> 2-methyl propan-2-ol  <span style="float: right;">(NCERT)</span></p> <p>(a) propanol                      (b) ethanone            (c) propanone                    (d) butane</p>
6	<p>What is the correct order of reactivity of alcohols in the following reaction?</p> $\text{R-OH} + \text{HCl} \xrightarrow{\text{ZnCl}_2} \text{R-Cl} + \text{H}_2\text{O}$ <p>(a) <math>1^\circ &gt; 2^\circ &gt; 3^\circ</math>                      (b) <math>1^\circ &lt; 2^\circ &gt; 3^\circ</math>            (c) <math>3^\circ &gt; 2^\circ &gt; 1^\circ</math>                      (d) <math>3^\circ &gt; 1^\circ &gt; 2^\circ</math></p>
7	<p>The major product of the following reaction is</p> <p style="text-align: center;">  <span style="float: right;">(2020)</span> </p> <p>(a)                       (b) </p> <p>(c)                       (d) </p>
8	<p>Reaction of tertiary butyl alcohol with hot Cu at <math>300^\circ\text{C}</math> produces</p> <p>(a) butanol            (b) butanal            (c) 2-butene            (d) 2-methylpropene</p>

9	<p>IUPAC name of <i>m</i>-cresol is</p>  <p>(a) 0-hydroxybenzoic acid  (b) 2-hydroxybenzoic acid  (c) Phenol-2-carboxylic acid  (d) 6-hydroxybenzoic acid</p>
10.	<p>Mark the correct order of decreasing acid strength of the following compounds.</p>  <p>(a) V &gt; IV &gt; II &gt; I &gt; III      (b) II &gt; IV &gt; I &gt; III &gt; V  (c) IV &gt; V &gt; III &gt; II &gt; I      (d) V &gt; IV &gt; III &gt; II &gt; I</p>

**Read the given passage and answer the questions that follow: (1 MARK EACH)**

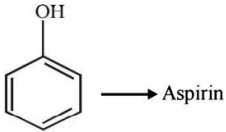
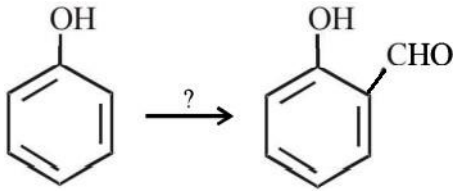
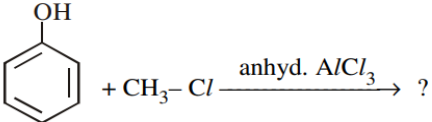
Methanol, CH<sub>3</sub>OH, also known as ‘wood spirit’, was produced by destructive distillation of wood. Today, most of the methanol is produced by catalytic hydrogenation of carbon monoxide at high pressure and temperature and in the presence of ZnO – Cr<sub>2</sub>O<sub>3</sub> catalyst.



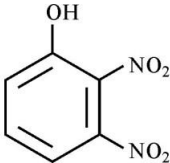
Methanol is a colourless liquid and boils at 337 K. It is highly poisonous in nature. Ingestion of even small quantities of methanol can cause blindness and large quantities causes even death. Methanol is used as a solvent in paints, varnishes and chiefly for making formaldehyde.

11. What happens when methanol is subjected to PCC?
12. Give two applications of methanol.
13. Mention the IUPAC name of the compound formed when methanol is subjected to thionyl chloride.
14. Give one commercial preparation of methanol.
15. Which is readily soluble in water – methanol or butan-1-ol.



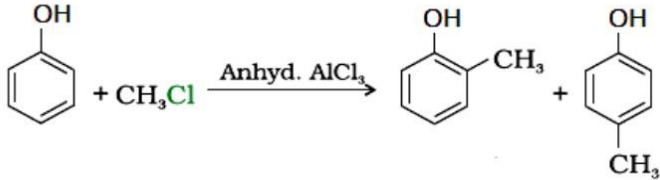
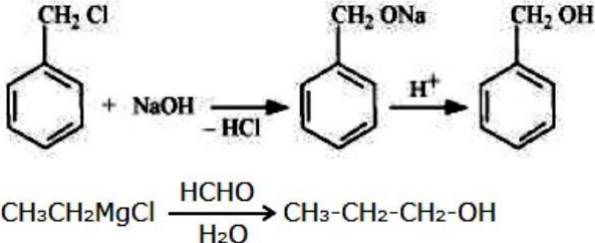
**Question – Answer Type: (Previous Years’ Board Questions)**

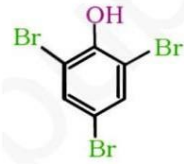
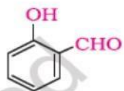
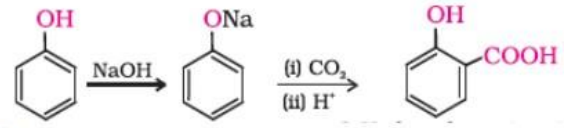
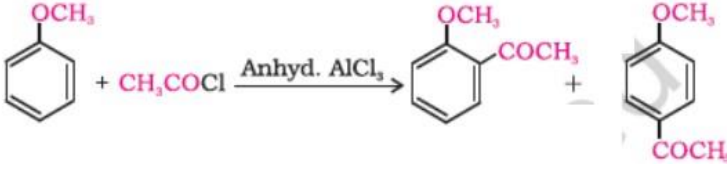
16	Convert: 	1
17	Give a reason: <b>Acid catalysed dehydration of t-butanol is faster than that of n-butanol.</b>	1
18	Give a chemical test to distinguish : Ethanol and Propan-2-ol	1
19	Complete the following: 	1
20	Complete the equation: $(\text{CH}_3)_3\text{CBr} + \text{NaOMe} \rightarrow$	1
21	Write the major product in the following equations : (i) $\text{CH}_3 - \text{CH}_2\text{OH} \xrightarrow{\text{PCl}_5} ?$ (ii) 	2

22	$\text{CH}_3 - \text{Cl} + \text{CH}_3\text{CH}_2 - \text{ONa} \rightarrow ?$ <p>a. Identify the name of the reaction. b. What are the products formed.</p>	2
23	<p>How are the following conversions carried out ?</p> <p>(i) Benzyl chloride to Benzyl alcohol (ii) Ethyl magnesium chloride to Propan-1-ol</p>	2
24	<p>Convert the following</p> <p>a. Propene to Propan-2-ol</p> <p>b. Propene to Propan-1-ol</p>	2
25	<p>Write the IUPAC name of the given compound :</p> $\text{HO} - \text{CH}_2 - \text{CH} = \underset{\begin{array}{c}   \\ \text{CH}_3 \end{array}}{\text{C}} - \text{CH}_3$ <p>Name the reagents used in the following reactions :</p> $\text{CH}_3 - \text{CHO} \xrightarrow{?} \underset{\begin{array}{c}   \\ \text{OH} \end{array}}{\text{CH}_3 - \text{CH}} - \text{CH}_3$	2
26	<p>Predict the products of the following reactions :</p> <p>(i) <math>\text{CH}_3 - \text{CH} = \text{CH}_2 \xrightarrow[\text{ii) } 3 \text{ H}_2\text{O}_2 / \text{OH}^-]{\text{i) } \text{B}_2\text{H}_6} ?</math></p> <p>(ii) <math>\text{C}_6\text{H}_5 - \text{OH} \xrightarrow{\text{Br}_2 (\text{aq})} ?</math></p> <p>(iii) <math>\text{CH}_3\text{CH}_2\text{OH} \xrightarrow{\text{Cu}/573 \text{ K}} ?</math></p>	3
27	<p>(a) Write the major product(s) in each of the following reactions :</p> <p>(i) <math>\text{CH}_3 - \underset{\begin{array}{c}   \\ \text{CH}_3 \end{array}}{\overset{\begin{array}{c} \text{CH}_3 \\   \end{array}}{\text{C}}} - \text{O} - \text{CH}_3 + \text{HI} \longrightarrow</math></p> <p>(ii) <math>\text{CH}_3 - \text{CH}_2 - \underset{\begin{array}{c}   \\ \text{OH} \end{array}}{\text{CH}} - \text{CH}_3 \xrightarrow{\text{Cu}/573 \text{ K}}</math></p> <p>(iii) <math>\text{C}_6\text{H}_5 - \text{OH} \xrightarrow[\text{(ii) } \text{H}^+]{\text{(i) } \text{CHCl}_3 + \text{aq} \cdot \text{NaOH}}</math></p>	3

28	<p>Write the chemical reaction involved in the following reactions :</p> <p>(i) Kolbe's reaction</p> <p>(ii) Friedal-Crafts acetylation of anisole</p> <p>Distinguish between :</p> <p>(i) Ethanol and phenol</p>	2, 1
29	<p>What happens when</p> <p>(i) phenol reacts with Bromine water ?</p> <p>(ii) ethanol reacts with <math>\text{CH}_3\text{COCl}</math>/pyridine ?</p> <p>(iii) anisole reacts with HI ?</p> <p>Write the chemical equations involved in the above reactions.</p>	3
30	<p>Write IUPAC name of the following compound :</p>  <p>Explain mechanism for hydration of acid catalyzed ethene :</p> $\text{CH}_2 = \text{CH}_2 + \text{H}_2\text{O} \xrightarrow{\text{H}^+} \text{CH}_3 - \text{CH}_2 - \text{OH}$	3

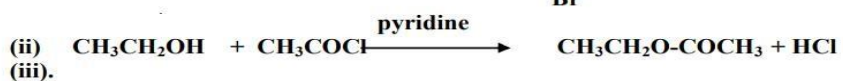
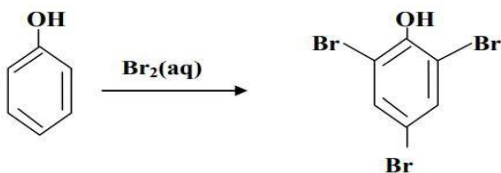
<b>ANSWER KEY</b>	
1	c
2	a
3	a
4	d
5	c
6	c
7	a
8	d
9	b
10	b
11	Methanal
12	paints, varnishes
13	Chloromethane
14	$\text{CO} + 2\text{H}_2 \xrightarrow[573-673 \text{ K}]{\text{ZnO-Cr}_2\text{O}_3, 200-300 \text{ atm}} \text{CH}_3\text{OH}$
15	methanol ( lower carbon chain)

16	 <p>Phenol <math>\xrightarrow{\text{NaOH}}</math> Sodium Phenoxide <math>\xrightarrow[\text{(ii) H}^+]{\text{(i) CO}_2}</math> 2-Hydroxybenzoic acid (Salicylic acid) <math>\xrightarrow[\text{CH}_3\text{COOH}]{\text{H}^+, (\text{CH}_3\text{CO})_2\text{O}}</math> Aspirin</p>
17	Hint: T. alcohols are more stable than p. alcohols
18.	Hint: Lucas test
19.	 <p>Phenol <math>\xrightarrow{\text{CHCl}_3 + \text{aq NaOH}}</math> Intermediate <math>\xrightarrow{\text{NaOH}}</math> Salicylaldehyde</p>
20.	$\text{CH}_3 - \underset{\text{CH}_3}{\overset{\text{CH}_3}{\text{C}}} - \text{Br} + \text{CH}_3\text{ONa} \xrightarrow{\text{Elimination}} \text{CH}_3 - \underset{\text{CH}_3}{\overset{\text{CH}_3}{\text{C}}} = \text{CH}_2 + \text{CH}_3\text{OH} + \text{NaBr}$ <p style="text-align: center;">Isobutylene</p>
21	<p>i) <math>\text{CH}_3\text{-CH}_2\text{OH} \xrightarrow{\text{PCl}_5} \text{CH}_3\text{CH}_2\text{Cl}</math></p> <p>ii) </p>
22	<p>a. Williamson Synthesis</p> <p>b. <math>\text{CH}_3\text{Cl} + \text{CH}_3\text{CH}_2\text{-ONa} \longrightarrow \text{CH}_3\text{CH}_2\text{-O-CH}_3 + \text{NaCl}</math></p>
23	 <p><math>\text{C}_6\text{H}_5\text{CH}_2\text{Cl} + \text{NaOH} \xrightarrow{-\text{HCl}} \text{C}_6\text{H}_5\text{CH}_2\text{ONa} \xrightarrow{\text{H}^+} \text{C}_6\text{H}_5\text{CH}_2\text{OH}</math></p> <p><math>\text{CH}_3\text{CH}_2\text{MgCl} \xrightarrow[\text{H}_2\text{O}]{\text{HCHO}} \text{CH}_3\text{-CH}_2\text{-CH}_2\text{-OH}</math></p>
24	<p>a.</p> $\text{CH}_3\text{CH}=\text{CH}_2 + \text{H}_2\text{O} \xrightleftharpoons{\text{H}^+} \text{CH}_3\text{-}\underset{\text{OH}}{\text{CH}}\text{-CH}_3$

	<p>b.</p> $\text{CH}_3-\text{CH}=\text{CH}_2 + (\text{BH}_3)_2 \longrightarrow \text{CH}_3-\text{CH}_2-\text{CH}_2\text{BH}_2$ <p style="text-align: center;">Propene</p> $\text{CH}_3-\text{CH}=\text{CH}_2 \longrightarrow (\text{CH}_3-\text{CH}_2-\text{CH}_2)_2\text{BH}$ $\downarrow \text{CH}_3-\text{CH}=\text{CH}_2$ $(\text{CH}_3-\text{CH}_2-\text{CH}_2)_3\text{B}$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH} + \text{B}(\text{OH})_3 \xleftarrow{\text{H}_2\text{O}_2/\text{OH}^-}$ <p style="text-align: center;">Propan-1-ol                      Boric acid</p>
25	<p>3-Methylbut-2-en-1-ol</p> <p><math>\text{CH}_3\text{MgBr} / \text{H}_3\text{O}^+</math></p>
26	<p>i) <math>\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}</math></p>  <p>ii)</p> <p>iii) <math>\text{CH}_3\text{CHO}</math></p>
27	<p>a)</p> <p>i) <math>(\text{CH}_3)_3\text{C-I} + \text{CH}_3\text{-OH}</math></p> <p>i) <math>\text{CH}_3\text{-CH}_2\text{-C(=O)-CH}_3</math></p> <p>ii)</p> 
28	<p>.i)</p>  <p>ii).</p>  <p>(i) Warm each compound with iodine and sodium hydroxide.</p> <p style="text-align: center;">Phenol : No yellow ppt formed</p> <p style="text-align: center;">Ethanol: Yellow ppt of Iodoform are formed.</p>



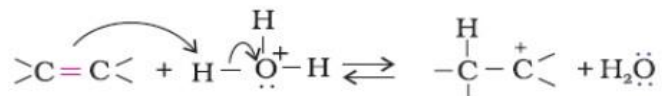
29

a).  
(i)

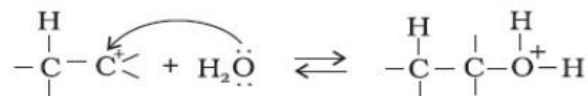
30

2,3-dinitro phenol

Protonation of alkene to form carbocation by electrophilic attack of  $\text{H}_3\text{O}^+$ .



Nucleophilic attack of water on carbocation.



Deprotonation to form an alcohol.

